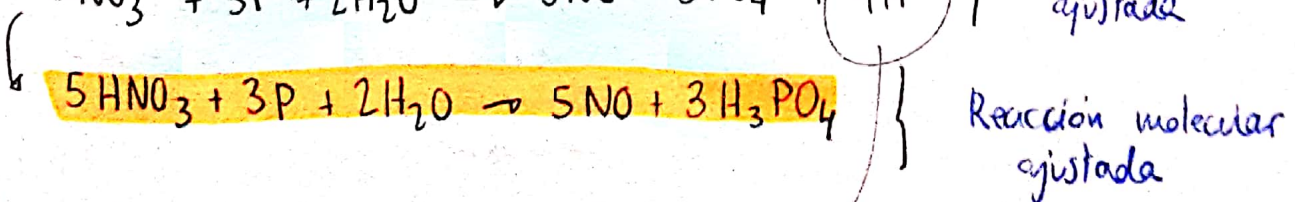
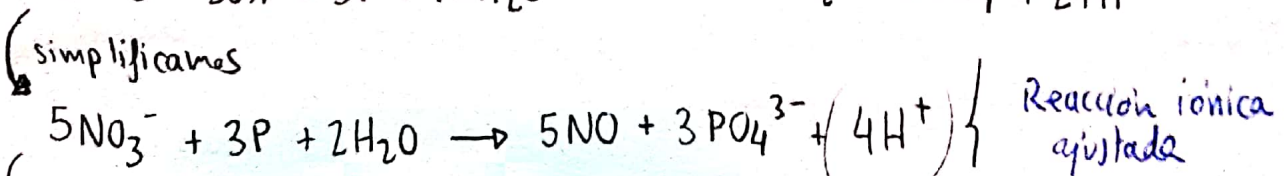
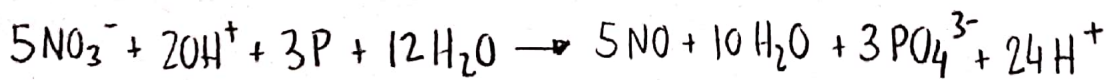
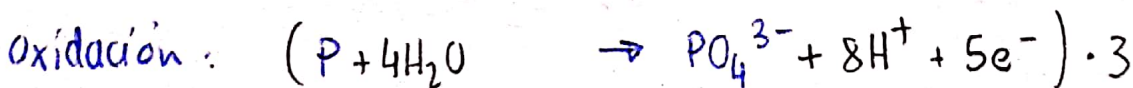
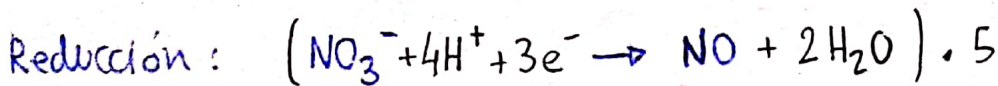
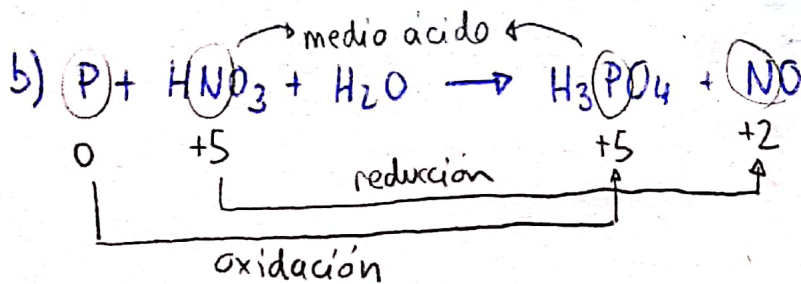
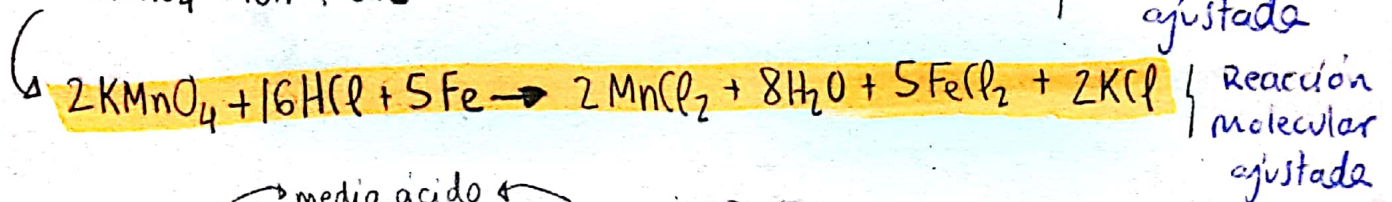
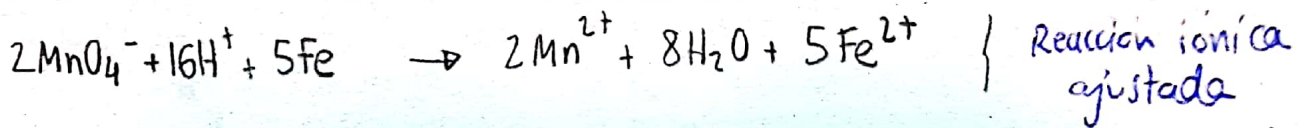
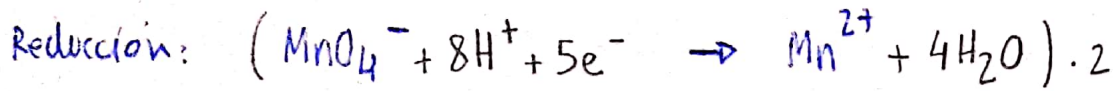
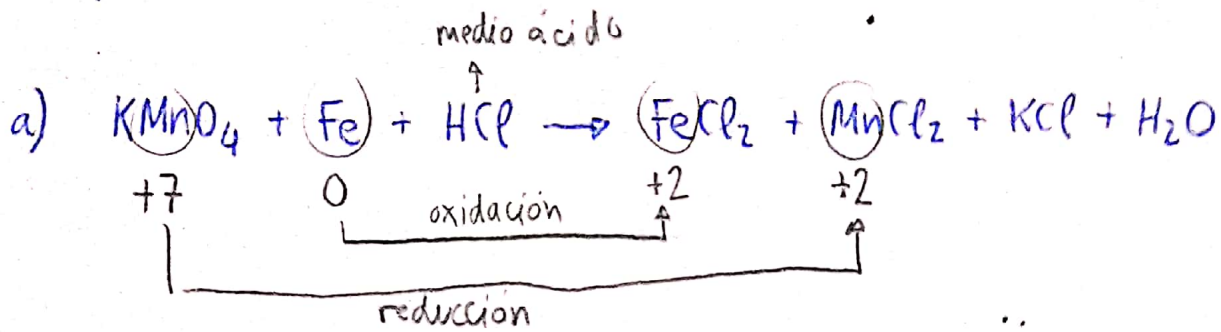
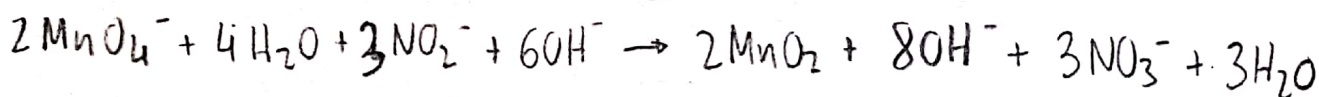
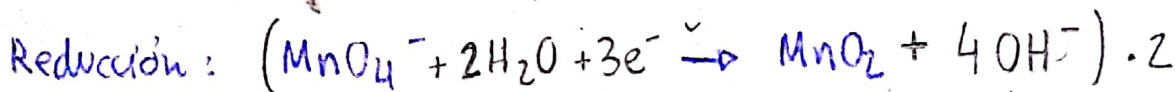
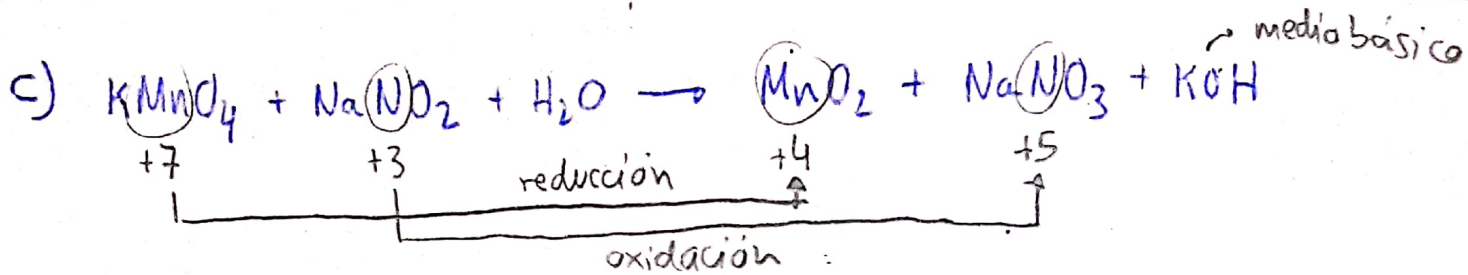


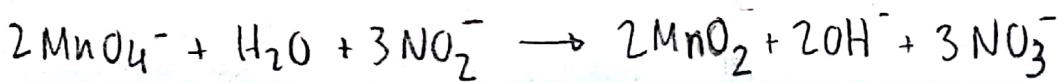
5) Ajusta las siguientes reacciones redox por el método del ion-electrón.



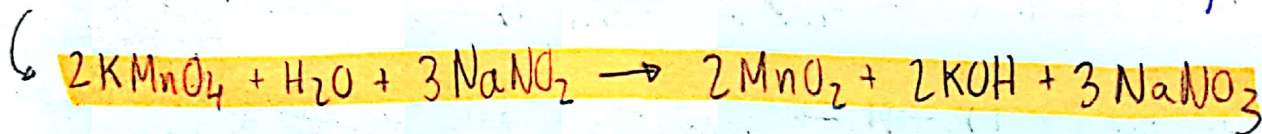
aunque nos indiquen 4H^+
hay que usar 9H^+ para
product $3\text{H}_3\text{PO}_4$



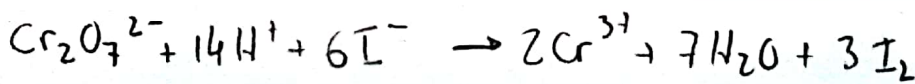
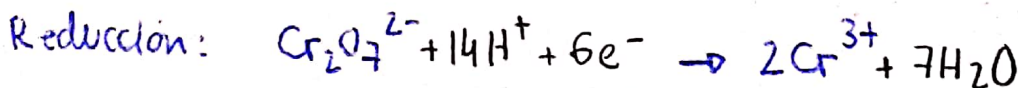
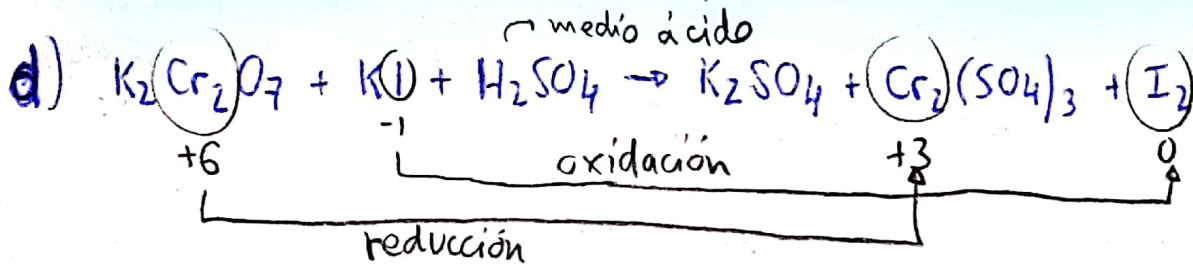
(simplificamos)



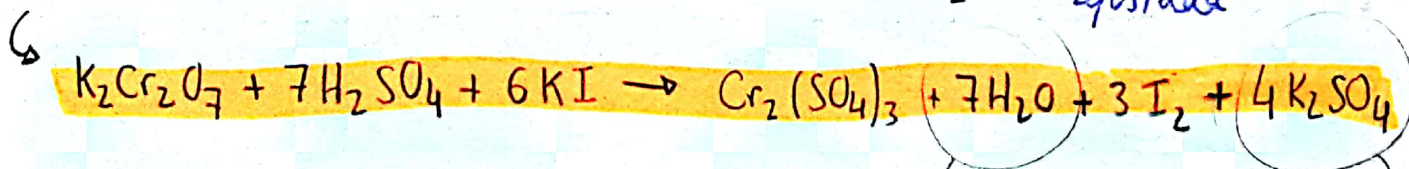
Reacción iónica ajustada



Reacción molecular ajustada



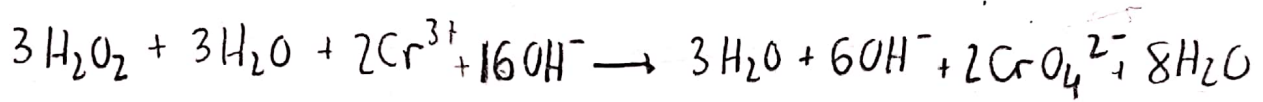
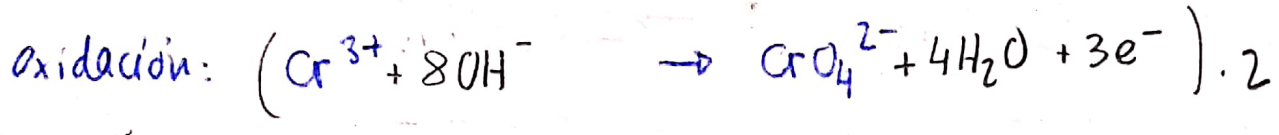
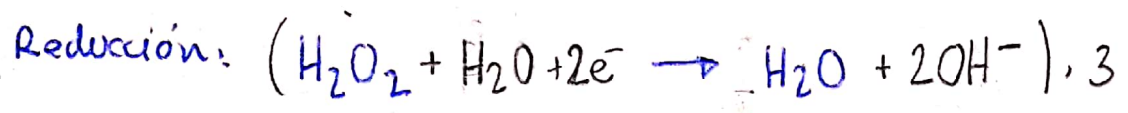
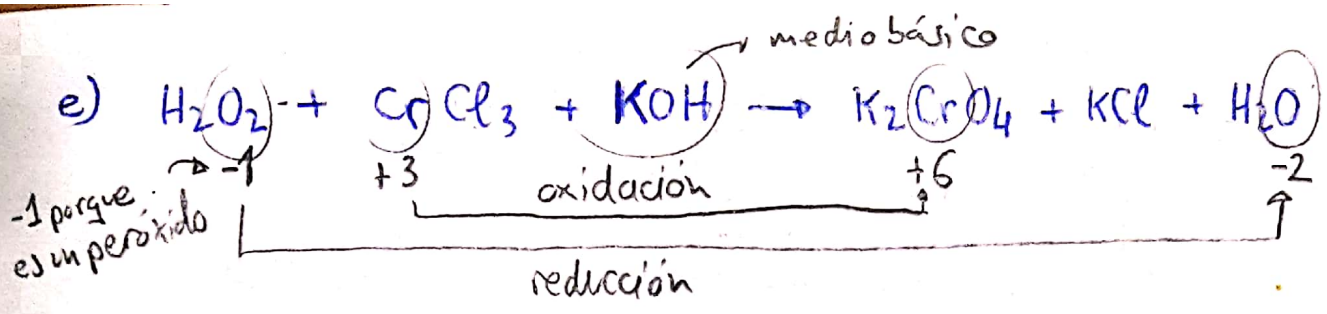
Reacción iónica ajustada



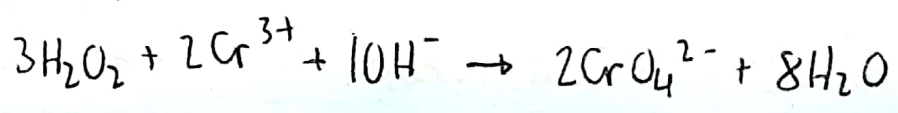
aunque no aparezca en el enunciado es necesaria

Reacción molecular ajustada

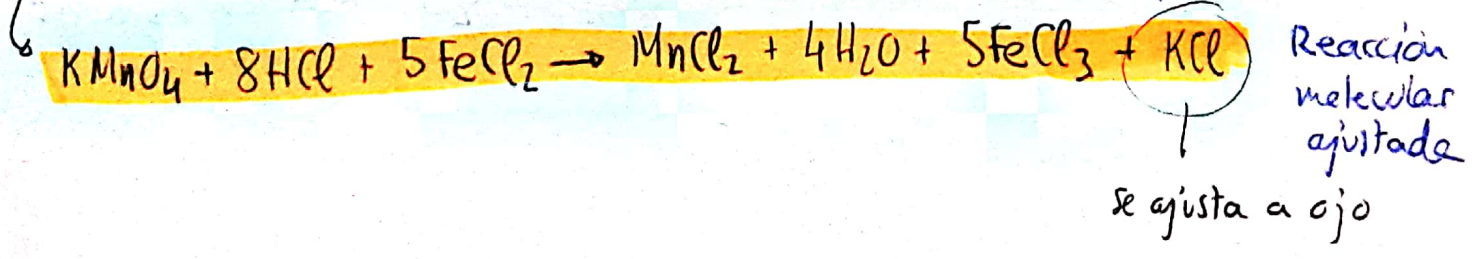
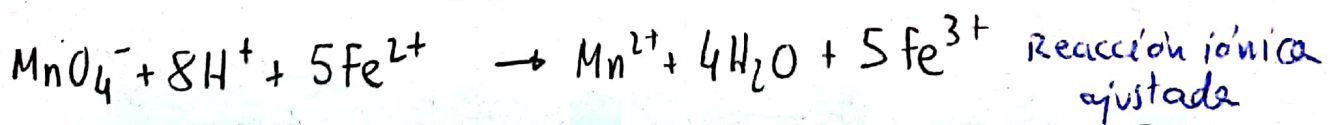
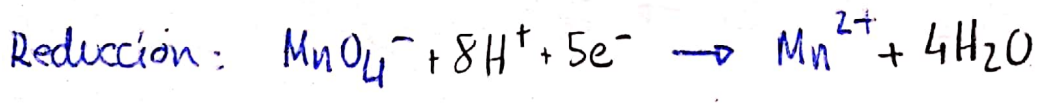
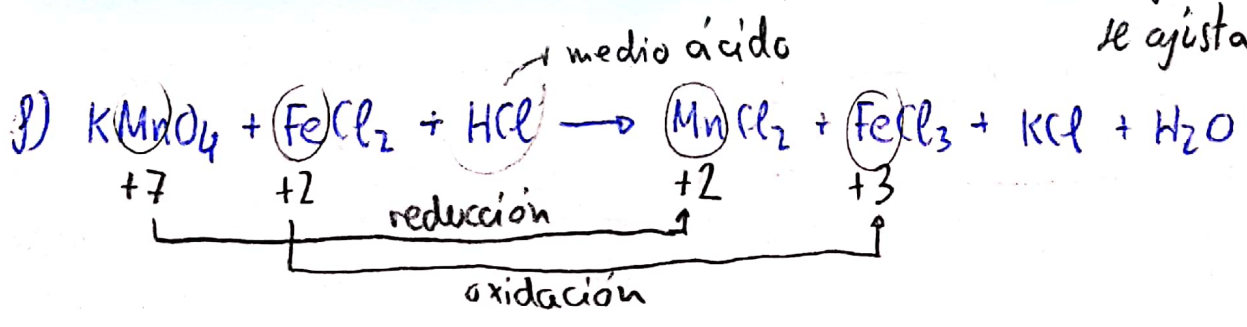
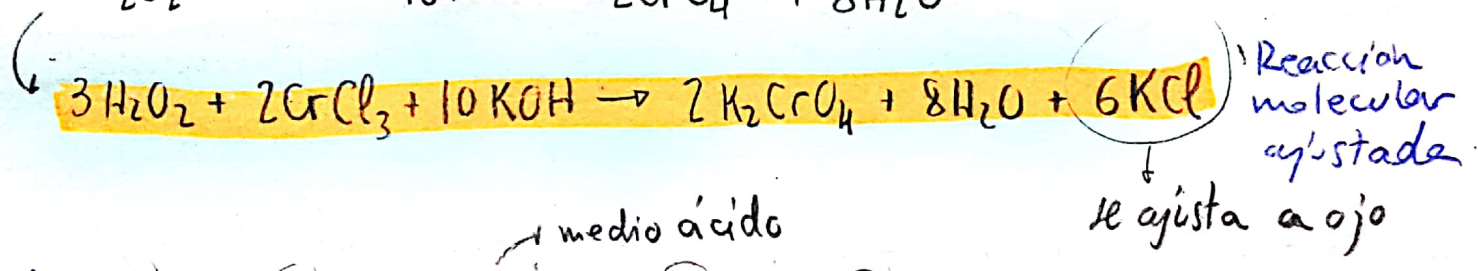
se ajusta a ojo

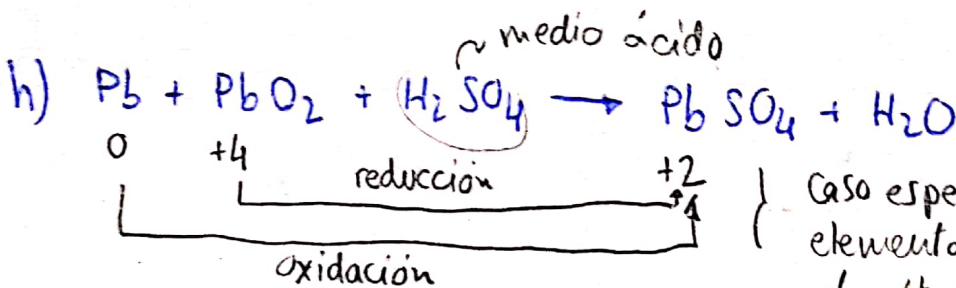
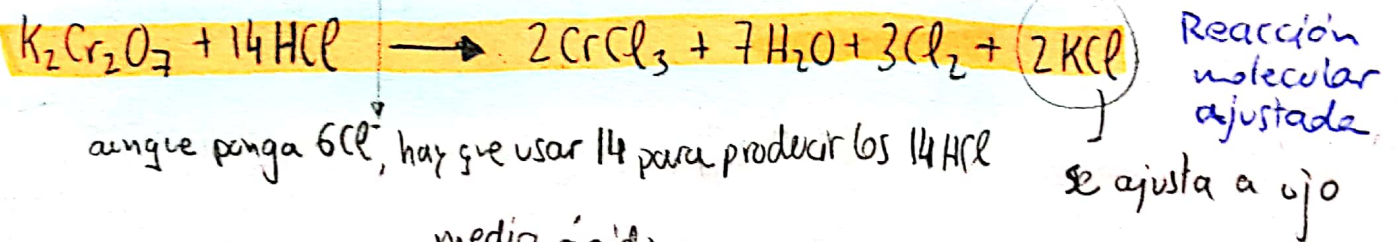
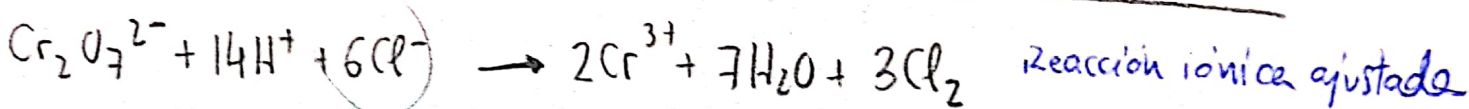
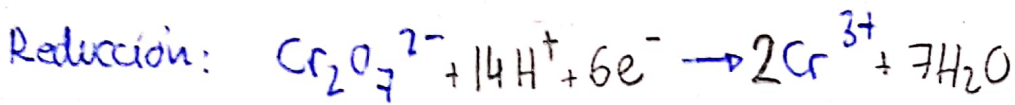
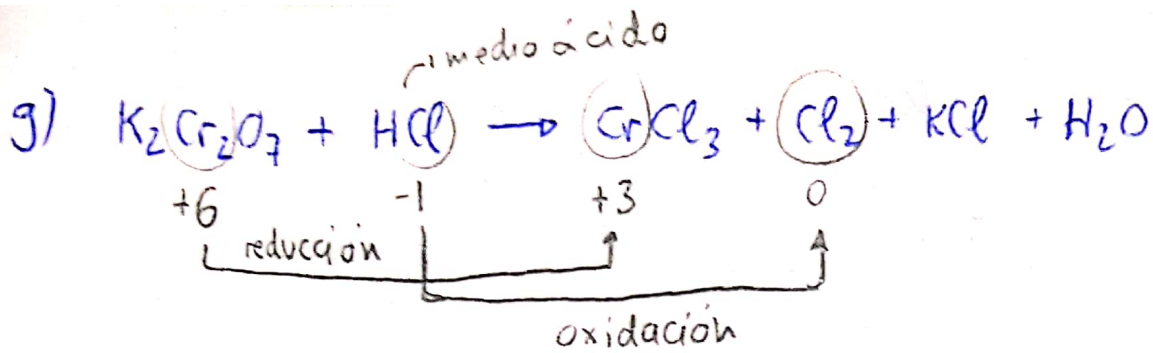


(simplificamos)

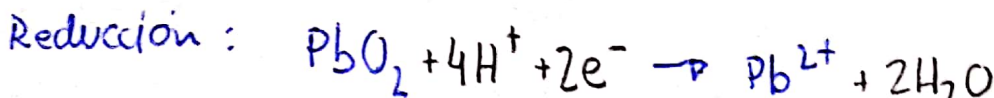


Reacción iónica ajustada





Caso especial en el que el mismo elemento se oxida y se reduce. Sin embargo, el método no cambia.



en este caso no es necesario multiplicar por nada.

